

Thermodynamics Multiple Choice

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Thermodynamics Multiple choice Questions :-1. Which of the following variables controls the physical properties of a perfect gas (a) pressure (b) temperature (c) volume (d) all of the above (e) atomic mass. Ans: d. 2. Which of the following laws is applicable for the behavior of a perfect gas (a) Boyle's law (b) Charles'law (c) Gay-Lussac law

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This set of Thermodynamics Multiple Choice Questions & Answers focuses on "Maximum Work in a Reversible Process". 1. For a process from state 1 to state 2, heat transfer in a reversible process is given by a) $Q_{\text{for reversible}} = (T_0) \cdot (S_1 - S_2)$ b) $Q_{\text{for reversible}} = (T_0) \cdot (S_2 - S_1)$ c) $Q_{\text{for reversible}} = (T_0) / (S_1 - S_2)$ d) $Q_{\text{for reversible}} = (T_0) / (S_2 - S_1)$ View ...

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a $2pv - 1/2 \cdot \#960v^2$ b $2pv + 1/2 \cdot \#960v^2$ c $2pv$ d None of the above. N rods of equal length L and equal cross-section area

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having thermal conductivity from k to nk respectively are joined together to form a rod of length NL .

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Thermodynamics Multiple Choice Questions and Answers

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This set of Thermodynamics Multiple Choice Questions & Answers (MCQs) focuses on "Available Energy Referred to a Cycle". 1. Which of the following is a type of energy? a) high grade energy b) low grade energy c) both of the mentioned d)

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none of the mentioned View Answer

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MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be ____ and has a ____ ΔH at constant pressure. A)endothermic, positive B)endothermic, negative C)exothermic, negative D)exothermic, positive E)exothermic, neutral

AP Chemistry Practice Test, Ch. 6: Thermochemistry ...

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Question 6 A drug used to treat hypertension undergoes a decomposition reaction to give an insoluble product. Calculate the temperature at which this reaction becomes spontaneous if the enthalpy of the reaction at 298 K is 51 kJ mol⁻¹ and the entropy of the reaction at this temperature is 118.74 J K⁻¹ mol⁻¹.

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